Biology, Castle View High School  
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**Analyzing Molecules**

**For each of the molecules below:**

1. **Circle all Hydrogens involved in polar covalent bonds, and which therefore have partial positive charge, and which can therefore form Hydrogen bonds.**
2. **Draw a Triangle all parts of the molecule to which a Hydrogen could form a Hydrogen Bond.**
3. **Draw a Box around Hydrogens (or groups of Hydrogens) with Non-Polar Covalent bonds, and which can therefore form Van Der Walls attractions.**

*Note: Some of molecules below take shortcuts to simplify the drawings:*

* *Groups of atoms may be shown together as a chemical formula where you need to determine how they are bound together*
* *Not all “extra” pairs of electrons are shown*
* *Where two lines are connected at a corner, or a line simply ends without indicating an atom, the corner or end is assumed to be a Carbon with as many hydrogens as it needs.*

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